

## The crystal structure of a third polymorph of $\text{Cu}_5(\text{PO}_4)_2(\text{OH})_4$

GERALD L. SHOEMAKER, JAMES B. ANDERSON AND EDWARD KOSTINER<sup>1</sup>

*Institute of Materials Science and Department of Chemistry  
University of Connecticut  
Storrs, Connecticut 06268*

### Abstract

A phase formed in the  $\text{CuO}-\text{P}_2\text{O}_5-\text{H}_2\text{O}$  system is an additional polymorph of pseudomalachite,  $\text{Cu}_5(\text{PO}_4)_2(\text{OH})_4$ . This phase crystallizes in the space group  $P\bar{1}$  with  $a = 4.445(1)$ ,  $b = 5.873(1)$ ,  $c = 8.668(3)\text{\AA}$ ,  $\alpha = 103.62(2)$ ,  $\beta = 90.35(2)$ ,  $\gamma = 93.02(1)^\circ$ . The structure, solved by Patterson methods, was refined by full-matrix least-squares techniques to a residual  $R = 0.022$  ( $R_w = 0.036$ ). The four crystallographically unique copper atoms lie in characteristically distorted (four short and two long bonds) six-coordinated sites which link together to form two-dimensional copper-containing sheets. As in the other two polymorphs, these sheets (which are joined together by phosphate tetrahedra and hydrogen bonds) can be derived from a two-dimensional framework of edge-sharing, copper-containing polyhedra.

### Introduction

The crystal structure of pseudomalachite (PM) was determined by Ghose (1963) and refined by Shoemaker *et al.* (1977). Recently, we described (Anderson *et al.*, 1977) the crystal structure of a synthetic polymorph of pseudomalachite which, as a convenience, we referred to as PPM.

During the course of a series of hydrothermal experiments designed to establish the conditions for equilibrium between PM and PPM, a careful visual inspection with a high-intensity light source revealed crystals which had a slightly bluer color than the surrounding products. X-ray powder diffraction examination revealed the presence of diffraction lines which could not be explained by the known polymorphs PM and PPM. A powder photograph of the blue-green material alone revealed it to be a third polymorph of  $\text{Cu}_5(\text{PO}_4)_2(\text{OH})_4$ , which we shall call QPM.

This paper presents the results of our refinement of the structure of QPM and a description of its structure in terms of polyhedral linkages within a two-dimensional net. The following paper (Shoemaker and Kostiner, 1981) will present a more detailed summary of polymorphism among compounds of the stoichiometry  $\text{Cu}_5(\text{PO}_4)_2(\text{OH})_4$ .

### Experimental

Once the identity of the new phase had been established, it was found among the products of a number of previous experiments in hydrothermal copper phosphate systems. Additional experiments were carried out in an attempt to determine the effects of temperature and pressure variation on the increasingly complex equilibrium between the polymorphs of  $\text{Cu}_5(\text{PO}_4)_2(\text{OH})_4$ .

QPM has never been obtained as the sole product of a hydrothermal reaction. Most commonly, it is found in close conjunction with PM, although in some experiments PPM was the only other polymorph observed. QPM has been grown by hydrothermal technique in sealed gold capsules from nutrient mixtures containing  $\text{Cu}_3(\text{PO}_4)_2$ , libethenite, malachite, or copper(II) nitrate with a variety of mineralizing solutions at pressures ranging from 1.2 to 4.0 kbar and bomb temperatures from 205 to 560°C. It has not been observed in mixtures whose pH after reaction was less than 6 or greater than 9.

The crystal used for data collection was obtained from the hydrothermal reaction between 67.4 mg  $\text{Cu}_3(\text{PO}_4)_2$  and 0.5 ml of approximately 1.5 M  $\text{K}_3\text{PO}_4$  solution held at 450°C and 3.8 kbar for four days. In this experiment, QPM formed dendritic crystals which were intimately mixed with PPM crystals and  $\text{CuO}$ . Significant amounts of QPM can be obtained by the

<sup>1</sup>Author to whom correspondence should be addressed.













H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL	H	K	L	FOBS	FCAL																						
4	4	1	5	32	32	4	4	4	4	3	443	444	5	3	1	38	39	5	1	5	154	155	5	-1	3	287	283	4	4	1	5	32	32	4	4	1	5	32	32	4	4	1	5	32	32	4	4	1	5	32	32										
4	4	2	3	149	150	4	4	4	4	3	101	100	5	3	3	290	290	5	1	6	98	98	5	-1	4	145	140	4	4	2	3	149	150	4	4	2	3	149	150	4	4	2	3	149	150	4	4	2	3	149	150										
4	4	2	4	45	54	4	4	4	4	5	91	94	5	3	4	288	289	5	0	7	36	23	5	-1	5	151	151	4	4	2	5	54	56	4	4	2	5	54	56	4	4	2	5	54	56	4	4	2	5	54	56										
4	4	2	5	54	56	4	5	4	5	4	190	190	5	2	6	170	175	5	0	6	425	424	5	-1	6	119	115	4	4	2	6	484	493	4	4	2	6	484	493	4	4	2	6	484	493	4	4	2	6	484	493										
4	4	2	8	116	116	4	5	2	132	127	5	2	3	280	285	5	0	4	75	77	5	-2	6	223	222	4	4	2	8	116	116	4	4	2	8	116	116	4	4	2	8	116	116	4	4	2	8	116	116												
4	4	3	5	63	66	4	5	1	45	44	5	2	2	283	284	5	0	3	114	111	5	-2	5	310	309	4	4	3	5	63	66	4	4	3	5	63	66	4	4	3	5	63	66	4	4	3	5	63	66												
4	4	3	4	65	67	4	6	0	452	453	5	2	1	124	123	5	0	2	461	463	5	-2	7	256	249	4	4	3	4	65	67	4	4	3	4	65	67	4	4	3	4	65	67	4	4	3	4	65	67												
4	4	3	3	402	400	4	6	1	98	96	5	2	0	493	494	5	0	1	123	124	5	-2	6	223	222	4	4	3	3	402	400	4	4	3	3	402	400	4	4	3	3	402	400	4	4	3	3	402	400												
4	4	3	2	260	263	4	2	112	117	5	1	0	169	163	5	0	0	196	200	5	-2	5	310	309	4	4	3	2	260	263	4	4	3	2	260	263	4	4	3	2	260	263	4	4	3	2	260	263													
4	4	0	386	387	5	4	1	90	96	5	1	2	41	27	5	-1	0	56	45	5	-2	4	122	117	4	4	0	386	387	5	4	1	3	192	195	5	1	3	192	195	5	1	4	233	227	5	-1	2	65	54	4	4	0	386	387	5	4	1	3	192	195
4	4	1	82	85	5	0	424	422	5	1	3	192	195	5	-1	1	263	269	5	-2	4	122	117	4	4	1	82	85	5	0	424	422	5	1	3	192	195	5	1	4	233	227	5	-1	2	65	54	4	4	1	82	85	5	0	424	422	5	1	3	192	195
2	2	258	257	5	3	0	136	134	5	1	4	233	227	5	-1	2	65	54	4	4	0	310	309	4	4	0	310	309	5	-2	4	122	117	4	4	0	310	309	5	-2	4	122	117	4	4	0	310	309	5	-2	4	122	117								